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Lab Sulfur Clock
Answers Lab Sulfur
Clock Answers The
Sulfur Clock - Reaction
Rate Lab Purpose: To
determine the
relationship between
reaction rate and
concentration

Background : $2\text{HCl} (\text{aq}) + \text{Na}_2\text{S}_2\text{O}_3 (\text{aq}) \rightleftharpoons 2\text{NaCl} (\text{aq}) + \text{SO}_2 (\text{aq}) + \text{S} (\text{s}) + \text{H}_2\text{O} (\text{aq})$
When sulfur is
produced in solution, it

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becomes insoluble.

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The order of reaction
with respect to the
iodate ion, m , must be
determined for the
following rate. It is
assumed that the order
of reaction with respect
to the bisulfate is zero,

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thus n is zero. Lab
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We will be able to measure reaction progression, because as the reaction occurs, solid Sulfur is being

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formed and 'falls out' of solution (a precipitate). Pre-Lab: The following data was collected by performing five separate trials of the reaction between Nitric Oxide and Oxygen gas.

$$2 \text{NO (g)} + \text{O}_2 \text{(g)} \rightarrow 2 \text{NO}_2 \text{(g)}$$

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THE SULFUR CLOCK REACTION. Purpose: To determine the effect of changing the concentration of a reactant on the rate of a. chemical reaction.

The reactants are 0.125 M $\text{Na}_2\text{S}_2\text{O}_3$ and 0.600 M HCl. The thiosulfate ion, $\text{S}_2\text{O}_3^{2-}$ (from the $\text{Na}_2\text{S}_2\text{O}_3$) decomposes in the presence of the .

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hydrogen ion, H⁺ (from
the HCl), in the
following reaction:

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The Sulfur Clock -
Reaction Rate Lab

Purpose: To determine
the relationship
between reaction rate
and concentration

Background : $2\text{HCl} (\text{aq}) + \text{Na}_2\text{S}_2\text{O}_3 (\text{aq}) \rightleftharpoons 2\text{NaCl} (\text{aq}) + \text{SO}_2 (\text{aq}) + \text{S} (\text{s}) + \text{H}_2\text{O} (\text{l})$

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Answers
aq) When sulfur is produced in solution, it becomes insoluble.

The Sulfur Clock - Reaction Rate Lab - Weebly

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The Sulfur Clock -
Reaction Rate Lab
Purpose: To determine
the relationship
between reaction rate
and concentration

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Background : $2\text{HCl} (\text{aq}) + \text{Na}_2\text{S}_2\text{O}_3 (\text{aq}) \rightleftharpoons 2\text{NaCl} (\text{aq}) + \text{SO}_2 (\text{aq}) + \text{S} (\text{s}) + \text{H}_2\text{O} (\text{aq})$ When sulfur is produced in solution, it becomes insoluble.

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Clock - Reaction Rate
Lab Purpose: To
determine the
relationship between

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reaction rate and
concentration

Background : $2\text{HCl} (\text{aq}) + \text{Na}_2\text{S}_2\text{O}_3 (\text{aq}) \rightleftharpoons 2\text{NaCl} (\text{aq}) + \text{SO}_2 (\text{aq}) + \text{S} (\text{s}) + \text{H}_2\text{O} (\text{aq})$
When sulfur is produced in solution, it becomes insoluble.

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I went to a
Thanksgiving dinner
with over 100 guests.
The chief was seen
coughing and not
wearing a mask.
Should I call the police

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on then?

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Iodine Clock Reaction
Lab Answers Part A:
Determining the
complete rate law The
order of reaction with
respect to the iodate
ion, m , must be
determined for the
following rate.

Iodine Clock Reaction Lab

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The Sulfur Clock Lab

Report Shady Sartaj

Partners: Hamad

Hayat, Lily Lou 21

February 2017 09

March 2017 SCH4U, 03

Submitted to: Mr. Lee

Introduction: Chemical

kinetics, which is the

area of chemistry that

deals with the rates of

reactions, is the

knowledge and study

of ways to make

chemical reactions

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occurs faster or slower.

Answers

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Reaction Rate Lab
Sulfur Clock. Chemistry
122 Reaction Rate Lab
Sulfur Clock 1.

INTRODUCTION. In past discussions you have dealt with the rate of a chemical reaction in a qualitative (descriptive) manner. During the present

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laboratory exercise we will view the rate of a chemical reaction quantitatively. Rate is taken to mean a change per unit ...

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THE IODINE "CLOCK"
REACTION 2015
www.proffeny.com 3

The appearance of the deep-blue complex tells us that at this point in time, (t color),

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sufficient I_2 has been produced by Reaction 1 to use up all of the $\text{S}_2\text{O}_3^{2-}$ (thiosulfate ion) originally added.

Experiment 1 The Iodine “Clock” Reaction

The Kinetics of the Iodine Clock Reaction
Pre-lab Assignment

Before coming to lab: •

Read the lab

thoroughly. • Answer

the pre-lab questions

that appear at the end

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of this lab exercise.

The questions should be answered on a separate (new) page of your lab notebook. Be sure to show all work, round answers, and include units on all answers.

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